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Answer Key A In any stoichiometry problem, the first step is always to calculate the number of moles of each reactant present. In this case, we are given the mass of $K_2Cr_2O_7$ in 1 mL of solution, which we can use to calculate the number of moles of K_2Cr ...

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Answer Key Chapter 12: Stoichiometry Mole Ratios Questions 1. Aluminum reacts with oxygen to produce aluminum oxide as follows:
 $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ a. If you use 2.3 moles

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of Al, how many moles of Al_2O_3 can you make? b. If you want 3.9 moles of Al_2O_3 , how many moles of O_2 are needed? 2. In the presence of sulfuric acid, metallic iron forms iron(III) sulfate: $2Fe + 3H_2SO_4 \rightarrow Fe_2(SO_4)_3 + 3H_2$...

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numerical relationships in chemical reactions
. What is stoichiometry based on? the law of
conservation of mass. What does
stoichiometry involved? balancing chemical
equations and mole ratios. mole ratio. a
conversion factor that relates the number of
moles of any two ...

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